

# Iodine



## **General Information**

### **Discovery**

Iodine was discovered by B. Courtois in 1811 in Paris, France.

### **Appearance**

Iodine is a blue-black, shiny solid which sublimes at room temperature into a blue-violet gas with an irritating odour.

### **Source**

Iodine occurs sparingly (0.05 parts per million) in sea-water. From this source it is assimilated by seaweeds, brines from old sea deposits and brackish waters from oil and salt wells.

Iodine is obtained commercially by extracting iodine vapour from processed brine, by ion exchange of brine or by liberating iodine from iodate obtained from nitrate ores.

### **Uses**

Iodine is used in several areas including pharmaceuticals, photographic chemicals, printing inks and dyes, catalysts and animal feeds.

### **Biological Role**

Iodine is an essential element, lack of which causes problems with the thyroid gland. The artificial radioisotope,  $^{131}\text{I}$ , with a half-life of 8 days, is used in treating the thyroid gland.

A solution of potassium iodide and iodine has germicidal effects and is used for the external treatment of wounds.

If iodine is in contact with the skin it can cause lesions, and iodine vapour is extremely irritating to the eyes and mucous membranes.

### **General Information**

Iodine forms compounds with many elements, but is less active than the other halogens. It dissolves readily in chloroform, carbon tetrachloride and carbon disulphide to form beautiful purple solutions. It is only sparingly soluble in water. Organic iodine compounds are important in organic chemistry.

## Physical Information

Atomic Number	53
Relative Atomic Mass ( $^{12}\text{C}=12.000$ )	126.9
Melting Point/K	386.7
Boiling Point/K	457.5
Density/kg m <sup>-3</sup>	4930 (293K)
Ground State Electron Configuration	[Kr]4d <sup>10</sup> 5s <sup>2</sup> 5p <sup>5</sup>
Electron Affinity (M-M <sup>-</sup> )/kJ mol <sup>-1</sup>	295

## Key Isotopes

Nuclide	<sup>123</sup> I	<sup>125</sup> I	<sup>127</sup> I	<sup>129</sup> I	<sup>131</sup> I
Atomic mass			126.9		
Natural abundance	0%	0%	100%	0%	0%
Half-life	13.3 h	60.2 days	stable	1.7x10 <sup>7</sup> yrs	8 days

## Ionisation Energies/kJ mol<sup>-1</sup>

M - M <sup>+</sup>	1008.4
M <sup>+</sup> - M <sup>2+</sup>	1845.9
M <sup>2+</sup> - M <sup>3+</sup>	3200
M <sup>3+</sup> - M <sup>4+</sup>	4100
M <sup>4+</sup> - M <sup>5+</sup>	5000
M <sup>5+</sup> - M <sup>6+</sup>	7400
M <sup>6+</sup> - M <sup>7+</sup>	8700
M <sup>7+</sup> - M <sup>8+</sup>	16400
M <sup>8+</sup> - M <sup>9+</sup>	19300
M <sup>9+</sup> - M <sup>10+</sup>	22100

## Other Information

Enthalpy of Fusion/kJ mol<sup>-1</sup> 15.27

Enthalpy of Vaporisation/kJ mol<sup>-1</sup> 41.67

### Oxidation States

Main I<sup>-1</sup>

Others I<sup>0</sup>, I<sup>III</sup>, I<sup>V</sup>, I<sup>VII</sup>

### Covalent Bonds/kJ mol<sup>-1</sup>

I - H 299

I - C 218

I - O 234

I - F 280

I - Cl 208

I - I 151

I - Si 234

I - P 184